

Assignment

1) how many grams of Zn are in 0.356 mole of Zn :

- A) 23 g B) 2.33 g C) 23.3 g D) 233 g

2) calculate the molar mass SO_2 :

- A) 54.07 B) 44.07 C) 74.07 D) 64.07

3) calculate the present compound by mass of H , P and O in the compound H_3PO_4 :

- A) 3.086% , 31.61% , 65.31% B) 31.61% , 3.086% , 65.31%
C) 31.61% , 65.31% , 3.086% D) 65.31% , 31.61% , 3.086%

4) it is compound of 40.92 percent carbon, 4.58 percent hydrogen, 54.50 percent oxygen

Determine it is empirical formula :

- A) $\text{C}_3\text{H}_4\text{O}_3$ B) CH_2O C) CHO D) $\text{C}_3\text{H}_2\text{O}_3$

5) empirical formula NO_2 , molar mass is compound= 90 g , determine the molecular formula :

- A) 2 B) NO C) N_2O_4 D) N_2O_2