

Assignment

1) Which of the following best defines the *atomic number* of an element :

A- The number of neutrons in the nucleus

B- The number of protons in the nucleus

C- The total number of protons and neutrons

D- The total number of electrons and neutrons

2) The *atomic number* of an element is equal to the number of :

A- Neutrons

B- Protons

C- Protons and neutrons

D- Electrons and
neutrons

3) The *mass number* of an atom is the total number of :

A- Protons only

B- Neutrons only

C- Electrons only

D- Protons and
neutrons

4) Which of the following pairs represents *isotopes* of the same element :

A- ^{12}C and ^{14}N

B- ^{14}C and ^{14}N

C- ^{12}C and ^{13}C

D- ^{13}C and ^{13}N

5) An atom has an atomic number of 8 and a mass number of 16, How many neutrons does this atom have :

A- 6

B- 8

C- 16

D- 24

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