Assignment

1) consider the following reaction : P 4 + 5O 2 → 2P 2 O 5 If 6.25 g of phosphorus is burned , what mass of oxygen :

A) 8.07 B) 4.5 C) 6.6 D) 10.0

2) calculate the mass of NH_3 produced when 5.60 g N_2 are reacted With 1.50 g H_2 according to the equation $N_2 + 3H_2 \rightarrow 2NH_3$:

A) 8.0 B) 6.8 C) 7.4 D) 4.3

3) A sample of 1.20 g of carbon was allowed to react with 1.60 g of Oxygen. If 1.65 g of CO_2 were produced, the percentage yield of CO_2 is :

A) 80% B) 70% C) 90% D) 75%

? / Urea [$(NH_2)_2$ CO] is prepared by reaction ammonia with carbon dioxide :

$2NH_3 + CO_2 \rightarrow (NH_2)_2$ CO + H_2O , in one process 637.2 g of NH_3 are treated with 1142 g of CO_2 :

4) which of the two reaction is the limiting reaction :

A) 37.4 B) 18.71 C) 25.95 D) 25

5) how much excess Reagan in grams is left at the end of the reaction :

A) 823.4 B) 1142 C) 654 D) 319

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CHEMISTRY